

Sulphur–Nitrogen Compounds, Part 3^a. Reaction of Heavy-metal Cations with Nitroso-disulphonate, Hydroxylamine-*N,N*-disulphonate, and Similar Anions

J. D. BIRCHALL and C. GLIDEWELL

Chemistry Department, University of St. Andrews, Fife KY16 9ST, Scotland, U.K.

Received September 14, 1977

Silver(I) ions are reduced to the metal by both $\text{ON}(\text{SO}_3)_2^-$ and $\text{HON}(\text{SO}_3)_2^-$, but react with $\text{HN}(\text{SO}_3\text{K})_2$ and $\text{K}_3\text{NS}_2\text{O}_6$ to yield $\text{Ag}_3\text{NS}_2\text{O}_6$ and $\text{Ag}_2\text{KNS}_2\text{O}_6$ respectively. Lead(II) and mercury(II), but not mercury(I), react with both $\text{ON}(\text{SO}_3)_2^-$ and $\text{HON}(\text{SO}_3)_2^-$ to yield MSO_4 ($M = \text{Pb}, \text{Hg}$) and mixtures of SO_4^{2-} and SO_3^{2-} . Stoichiometries are established, and possible mechanisms discussed, for some of these decompositions.

Introduction

As part of a study of sulphonylhydroxylamines [1, 2], we have sought routes to neutral, soluble analogues of Fremy's radical-dianion, nitrosyl disulphonate, $\text{ON}(\text{SO}_3)_2^-$ [3], of type $\text{ON}(\text{SO}_2\text{OR})$ ($R = \text{alkyl, aryl, etc.}$). One possible route to such species is the reaction of molecular halides with the heavy metal derivatives both of $\text{ON}(\text{SO}_3)_2^-$ and of hydroxylamine-*N,N*-disulphonate $\text{HON}(\text{SO}_3)_2^-$, oxidation of species $\text{HON}(\text{SO}_2\text{OR})_2$ being expected to yield $\text{ON}(\text{SO}_2\text{OR})_2$. In this paper we report the results of an investigation of the reactions of the heavy metal ions Ag^+ , Hg^{+2} , Hg_2^{+2} , Ti^+ and Pb^{+2} with the anions $\text{ON}(\text{SO}_3)_2^-$ and $\text{HON}(\text{SO}_3)_2^-$, and also with the closely analogous anions imidodisulphonate, $\text{HN}(\text{SO}_3)_2^-$, nitrilotrisulphonate $\text{N}(\text{SO}_3)_3^{3-}$, hydroxylamine-*N,O*-disulphonate $[(\text{O}_3\text{S})\text{ONH}(\text{SO}_3)]^{2-}$ and hydroxylamine-*N,N,O*-trisulphonate $[(\text{O}_3\text{S})\text{ON}(\text{SO}_3)_2]^{3-}$.

Salts with the metals of Groups I and II have been described for the anions $\text{ON}(\text{SO}_3)_2^-$ [3, 4], $\text{HON}(\text{SO}_3)_2^-$ [5, 6], $\text{HN}(\text{SO}_3)_2^-$ [5, 7, 8], $\text{N}(\text{SO}_3)_3^{3-}$ [3, 9], and $[(\text{O}_3\text{S})\text{ON}(\text{SO}_3)_2]^{2-}$ [3, 10], but no heavy metal derivatives have been reported for these anions except $\text{HN}(\text{SO}_3)_2^-$ [5] of which a number of silver(I), lead(II) and mercury(II) derivatives have been described, and $\text{HON}(\text{SO}_3)_2^-$ of which some

lead(II) derivatives have been reported [6]: all of these materials are of complex stoichiometry and unknown constitution. Of the $\text{HN}(\text{SO}_3)_2^-$ derivatives, those of silver include $\text{Ag}_3\text{NS}_2\text{O}_6$ [5, 7] and $\text{Ag}_2\text{NaS}_2\text{O}_6$ [5]; amongst the lead derivatives, no binary species such as $\text{Pb}_3(\text{NS}_2\text{O}_6)_2$ or $\text{HN}(\text{SO}_3)_2\text{Pb}$ appear to exist [5] but only hydroxylead species such as $(\text{HO-Pb})_3\text{NS}_2\text{O}_6$ and $[(\text{HOPb})_3\text{NS}_2\text{O}_6] \cdot \text{Pb}(\text{OH})_2$ [5]. Similarly no binary mercury(II) derivatives appear to have been described, but a large number of mixed derivatives, such as $\text{Hg}[\text{N}(\text{SO}_3\text{K})_2]_2$ exist [5, 7], together with a number of oxomercury species of bewildering composition, such as $\text{HNS}_2\text{O}_6 \cdot \text{Hg}_3\text{O}_2$ [5] and $(\text{NS}_2\text{O}_6)_2\text{Hg}_5\text{O}_2$ [7]. No binary heavy-metal derivatives of the $\text{HON}(\text{SO}_3)_2^-$ ion appear to exist [6]; only hydroxylead species, such as $(\text{HOPb})_3(\text{ONS}_2\text{O}_6)$ and $(\text{HOPb})_2(\text{HONS}_2\text{O}_6)$ have been described: these form part of a series of materials [6] in which the $\text{Pb}:\text{N}$ ratio varies from 3:1 to 2:3. The reported members of this series may be distinct compounds, or mixtures, or may be phases of apparently stoichiometric composition.

Experimental

Standard methods were employed for the preparation of $\text{ON}(\text{SO}_3\text{K})_2$ [4], $\text{HON}(\text{SO}_3\text{K})_2$ [11], $\text{HN}(\text{SO}_3\text{K})_2$ [11], $\text{N}(\text{SO}_3\text{K})_3$ [11], $(\text{KO}_3\text{S})\text{ONH}(\text{SO}_3\text{K})$ [11] and $(\text{KO}_3\text{S})\text{ON}(\text{SO}_3\text{K})_2$ [11]. Nitrates of Ag(I), Hg(II), Hg(I), Ti(I) and Pb(II) were AnalaR where available and all were used as supplied. $\text{KN}(\text{SO}_3\text{K})_2$ was prepared by reaction of $\text{HN}(\text{SO}_3\text{K})_2$ with aqueous KOH.

Analysis

Metals were determined using a Perkin-Elmer model 360 Atomic Absorption Spectrophotometer; $\text{ON}(\text{SO}_3)_2^-$ was determined iodimetrically, and SO_3^{2-} and SO_4^{2-} were determined gravimetrically. Hydrogen and nitrogen analyses were by the microanalytical laboratory of this Department.

^aPart 2, ref [2].

*Reactions of ON(SO₃K)₂**With Ag⁺*

Solutions of ON(SO₃K)₂ (23.95 mmol) and AgNO₃ (11.97 mmol) were mixed: the mixture became colourless after 1.5 h, and after 2 d a grey precipitate (1.2903 g: 11.96 mmol: 100%) was formed (Found: Ag, 98.2%). The filtrate contained sulphate(VI) (45.2 mmol: 94.4%).

With Hg⁺²

Solutions of ON(SO₃K)₂ (23.95 mmol) and Hg(NO₃)₂ (11.98 mmol) were mixed: the mixture became colourless within 1 h. A white precipitate (3.09 g) was formed (Found: Hg, 67.9; S, 11.2; N, 0.0%; HgO₄S requires: Hg, 67.6, S, 10.8%). The filtrate contained SO₄²⁻ (28.7 mmol) and SO₃²⁻ (5.4 mmol), mol ratio SO₄²⁻: SO₃²⁻ = 5.31:1.

With Pb⁺²

Solutions of ON(SO₃K)₂ (23.95 mmol) and Pb(NO₃)₂ (11.97 mmol) were mixed: the mixture became colourless within 1 h. A white precipitate (3.461 g) was formed (Found: Pb, 65.2; S, 10.2; N, 0.0%; O₄PbS requires: Pb, 68.3; S, 10.6%). The filtrate contained SO₄²⁻ (27.2 mmol) and SO₃²⁻ (5.82 mmol), mol ratio SO₄²⁻: SO₃²⁻ = 4.68:1.

With Tl⁺

Mixtures of ON(SO₃K)₂ and TlNO₃ in molar ratio 1:1, 2:1, 3:1 and 4:1 all became colourless within 12 h, but no insoluble products resulted.

*Reactions of HON(SO₃K)₂**With Ag⁺*

Solutions of HON(SO₃K)₂ (10.09 mmol) and AgNO₃ (10.09 mmol) were mixed, and kept in the absence of light for a week, during which time gases were evolved. A dark precipitate (1.0073 g: 9.33 mmol: 92.5%) was produced (Found: Ag, 98.2; S, 0.0%). The filtrate contained sulphate(VI) (19.0 mmol, 94.2%).

With Hg⁺²

Solutions of HON(SO₃K)₂ (1.00 mmol) and Hg(NO₃)₂ (2.00 mmol) were mixed together at pH 6.8 (phosphate buffer). A white precipitate (0.3815 g) was formed (Found: Hg, 67.9; S, 11.2; N, 0.0%; HgO₄S requires: Hg, 67.6; S, 10.8%). The filtrate contained mercury(II), (1.40 mmol), SO₄²⁻ (0.50 mmol) and SO₃²⁻ (0.50 mmol).

With Pb⁺²

Solutions of HON(SO₃K)₂ (9.71 mmol) and Pb(NO₃)₂ (9.71 mmol) were mixed together and after stirring during 24 h afforded a white solid (2.834 g: 9.35 mmol: 96.3%) (Found: Pb, 68.8; S, 11.1; N,

0.0%; O₄PbS requires: Pb, 68.3; S, 10.6%). The filtrate contained SO₄²⁻ (5.63 mmol) and SO₃²⁻ (4.02 mmol).

With Hg⁺² and Tl⁺

No reaction was observed between HON(SO₃)₂⁻ and either Hg⁺² or Tl⁺.

*Reactions of HN(SO₃K)₂**With Ag⁺*

Although no reaction was observed between equimolar solutions of HN(SO₃K)₂ and AgNO₃ during 3 weeks, when the molar ratio HN(SO₃K)₂: AgNO₃ was 1:3, yields of up to 68% of Ag₃NS₂O₆ were obtained (Found: Ag, 64.6; N, 2.6; S, 13.3%. Ag₃-NO₆S₂ requires: Ag, 65.1; N, 2.8; S, 12.9%).

With Pb⁺²

Addition of a solution of lead(II) acetate or nitrate (0.01 mol) to a stirred solution of HN(SO₃K)₂ (0.04 mol) yielded a white flocculent precipitate of (HOPb)N[SO₃(PbOH)]₂ (Found: H, 0.3; N, 1.4. Pb, 74.7; S, 7.2%. H₃NO₉Pb₃S₂ requires: H, 0.4; N, 1.6; Pb, 73.4; S, 7.6%).

With Hg⁺², Hg⁺², and Tl⁺

No reactions were observed between HN(SO₃K)₂ and Hg⁺², Hg⁺², and Tl⁺.

Reaction of KN(SO₃K)₂ with Silver Nitrate

An aqueous solution of KN(SO₃K)₂ (0.60 g: 2.06 mmol) was slowly stirred into a solution of AgNO₃ (0.74 g: 4.10 mmol). A pale sand-coloured precipitate of Ag₂KNS₂O₆ was formed. Found: Ag, 49.7; K, 9.1; N, 3.3; S, 14.6%. Ag₂KNO₆S₂ requires: Ag, 50.3; K, 9.1; N, 3.3; S, 14.9%.

Reactions of N(SO₃K)₃

No reactions were observed between N(SO₃K)₃ and AgNO₃, Hg(NO₃)₂, Hg₂(NO₃)₂, or TlNO₃. When solutions of N(SO₃K)₃ (8.08 mmol) and Pb(NO₃)₂ (7.97 mmol) were mixed, an immediate white precipitate (7.83 mmol, 98.2%) of PbSO₄ was formed. Found: Pb, 68.5; S, 10.8%. O₄PbS requires: Pb, 68.3; S, 10.6%.

Reactions of (KO₃S)ONH(SO₃K) and (KO₃S)ON(SO₃K)₂

No reactions were observed between either of these salts, and any of the metal ions Ag⁺, Hg⁺², Hg⁺², Tl⁺ and Pb⁺².

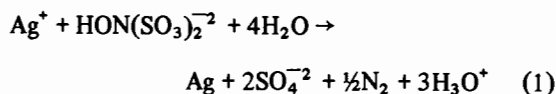
Reactions with Molecular Halides

No reaction was observed in, and the starting materials were recovered unchanged from, the following systems:

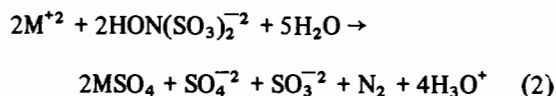
- (a) PhCH₂Br + ON(SO₃K)₂ in CH₃CN
(reflux, 3 h)
- (b) PhCH₂Br + HON(SO₃K)₂ in CH₃CN
(reflux, 2 h)
- (c) EtI + Ag₂KNS₂O₆ in C₆H₆
(R.T., 48 h)
- (d) EtI + Ag₃NS₂O₆ in C₆H₆
(R.T., 48 h)
- (e) C₈H₁₇Br + Ag₃NS₂O₆ in C₆H₆
(reflux, 48 h)
- (f) Me₃SiCl + HON(SO₃K)₂ in CH₃CN or
(CH₃)₂SO (R.T., 48 h)
- (g) Me₃SiCl + HN(SO₃K)₂ in CH₃CN or
(CH₃)₂SO (R.T., 48 h)

Results and Discussion

With silver nitrate, the HON(SO₃)₂⁻² ion is converted quantitatively to sulphate(VI) with concurrent reduction of Ag(I) to metallic silver, equation (1):

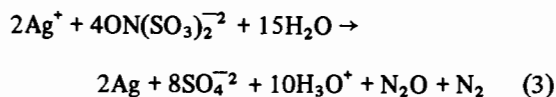


By contrast, both Pb(II) and Hg(II) nitrates yield precipitate of MSO₄ (M = Pb, Hg), leaving equimolar sulphate(VI) and sulphate(IV) in solution, equation (2):

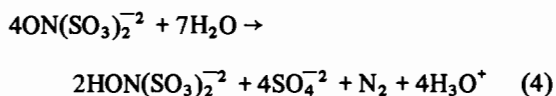


This reaction of lead(II) nitrate should be contrasted with that of basic lead(II) acetate, Pb(OCOCH₃)₂·nPb(OH)₂, which yields the hydroxolead(II) hydroxylamine disulphonate (HOPb)₃(ONS₂O₆) [6]. Neither of Hg⁺² and Tl⁺ appears to react (at least when as the nitrate) with the HON(SO₃)₂⁻² ion.

Silver ions react with the ON(SO₃)₂⁻² radical-ion in a reaction of stoichiometry 1Ag:2ON(SO₃)₂⁻², to yield metallic silver and sulphate(VI), equation (3):



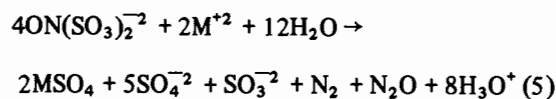
it is possible that the reaction proceeds *via* the HON(SO₃)₂⁻² ion, equation (4) [12]:



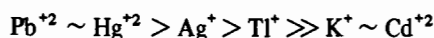
this ion then would react further according to equation (1). However it should be noted that reaction (3) is complete within 1.5 hours, although the uncatalysed

decomposition of ON(SO₃)₂⁻², equation (4) is slow, taking *ca.* 500 h for completion at pH 8.

Fast decomposition of ON(SO₃)₂⁻² also occurs in the presence of Pb⁺² and Hg⁺², to yield MSO₄ (M = Pb, Hg) together with sulphate(VI) and sulphate(IV) in a molar ratio of 5:1, equation (5):



Equation (5) can be regarded as the sum of equations (2) and (4), just as for Ag(I), equation (3) could be regarded as the sum of equations (1) and (4). The rate of decomposition of ON(SO₃)₂⁻² is found to depend on the metal ion present, thus:



The rate of decomposition of ON(SO₃)₂⁻² has been found [12] to be pH dependent, being least at pH 8, and to be catalysed also by nitrate(III) and chloride [4, 12], but these are both absent from the systems in this study. The production of a water-insoluble product is not a necessary requirement for a fast rate of decomposition, but it is perhaps significant that each of the metals found here to promote fast decomposition of ON(SO₃)₂⁻² has at least two oxidation states accessible (albeit only transiently in some instances) in aqueous solution, so that one- or two-electron redox reactions involving the metal are possible; by contrast, with cadmium which has but a single oxidation state accessible in aqueous solution, the rate of decomposition of ON(SO₃)₂⁻² is essentially unchanged from that of the pure potassium salt, and no insoluble product results. The hypothesis that at least two accessible oxidation states of the cation are necessary for fast decomposition rationalises adequately the observation that derivatives of ON(SO₃)₂⁻² have hitherto been described only of Group I and Group II metals, and further, suggests that other simple stable derivatives will be accessible only of metal ions such as Sc⁺³, Y⁺³, La⁺³, Zn⁺², and Cd⁺².

Silver ions react, in the presence of potassium ions, with both HN(SO₃)₂⁻² and N(SO₃)₂⁻³, although not with N(SO₃)₃⁻³ to yield respectively Ag₃NS₂O₆ and Ag₂KNS₂O₆: the constitutions of these compounds are probably complex. A similar precipitation without decomposition results from the reaction of lead(II) nitrate with HN(SO₃)₂⁻² to yield the previously described [6] hydroxolead species (HO-Pb)₃(NS₂O₆): that lead nitrate produces an hydroxo compound with HN(SO₃)₂⁻² is surprising since with HON(SO₃)₂⁻² the basic acetate was required for the production of (HOPb)₃(ONS₂O₆) [6], the nitrate yielding only PbSO₄.

The only reaction observed for N(SO₃)₃⁻³ was that with Pb⁺² ions, in which PbSO₄ is formed quantitatively

vely. In contrast to the reactivity of hydroxylamine-*N,N*-disulphonate, $\text{HON}(\text{SO}_3)_2^-$, its isomer, the *N,O*-disulphonate, $[(\text{O}_3\text{S})\text{ONH}(\text{SO}_3)]^-$ reacts with none of the metal ions studied here; similarly the *N,N,O*-trisulphonate $[(\text{O}_3\text{S})\text{ON}(\text{SO}_3)_2]^-$ shows no reaction with these metal ions.

The salt $\text{HON}(\text{SO}_3\text{K})_2$ was found to react neither with PhCH_2Br nor with Me_3SiCl : similarly no reaction was found between $\text{ON}(\text{SO}_3\text{K})_2$ and PhCH_2Br , or between Me_3SiCl and $\text{HN}(\text{SO}_3\text{K})_2$. It was the failure of the potassium salts to react with molecular halides which led us to attempt the preparation of heavy-metal derivatives: however these were prepared only from the anion $\text{HN}(\text{SO}_3)_2^-$; neither $\text{Ag}_3\text{NS}_2\text{O}_6$ nor $\text{Ag}_2\text{KNS}_2\text{O}_6$ was found to react with haloalkanes even during 48 hours reflux.

References

- 1 J. D. Birchall and C. Glidewell, *J. Chem. Soc. Dalton*, 10 (1977).
- 2 J. D. Birchall and C. Glidewell, *J. Chem. Soc. Dalton*, 1976 (1977).
- 3 E. Frémy, *Ann. Phys. Chim.*, 15, 408 (1845).
- 4 W. Moser and R. A. Howie, *J. Chem. Soc. A*, 3039 (1968).
- 5 E. Divers and T. Haga, *J. Chem. Soc.*, 61, 943 (1892).
- 6 E. Divers and T. Haga, *J. Chem. Soc.*, 65, 523 (1894).
- 7 E. Berglund, *Ber.*, 9, 252, 1896 (1876).
- 8 M. Woronin, *Russ. J. Chem.*, 3, 273 (1859).
- 9 F. Raschig, *Ann.*, 152, 336 (1869).
- 10 T. Haga, *J. Chem. Soc.*, 85, 78 (1904).
- 11 W. G. Palmer, "Experimental Inorganic Chemistry", Cambridge University Press (1954).
- 12 J. H. Murib and D. M. Ritter, *J. Am. Chem. Soc.*, 74, 3394 (1952).